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Mastering Chemistry - Chapter 4. For soluble salts, the molarity of a particular ion is equal to the molarity of that compound times the subscript for that ion. For example, 1 M of  $\text{AlCl}_3$  is 1 M in  $\text{Al}^{3+}$  and 3 M in  $\text{Cl}^-$ . 1 M of  $(\text{NH}_4)_2\text{SO}_4$  is 2 M in  $\text{NH}_4^+$  and 1 M in  $\text{SO}_4^{2-}$ .

### Answers To Mastering Chemistry Chapter 4

Step 4 of 6 Sodium sulfide : When dissolved in water it forms hydroxide ions therefore, it is alkaline in nature. The presence of sulfide ion can be confirmed by the following test.

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Mastering Chemistry Chapter 4  $\text{Zn (s)} + \text{MnI}_2 \text{ (aq)} \rightarrow \text{ZnI}_2 \text{ (aq)} + \text{Mn (s)}$   $3\text{Hg (l)} + 2\text{Cr (NO}_3)_3 \text{ (aq)} \rightarrow 3\text{Hg (NO}_3)_2 + 2\text{Cr (s)}$   $3\text{FeBr}_2 \text{ (aq)} + 2\text{Au (s)} \rightarrow 3\text{Fe (s)} + 2\text{AuBr}_3 \text{ (aq)}$   $2\text{AgNO}_3 \text{ (aq)} + \text{Pb (s)} \rightarrow 2\text{Ag (s)} + \text{Pb (NO}_3)_2 \text{ (aq)}$   $\text{SnCl}_2 \text{ (aq)} + \text{Cu (s)} \rightarrow \text{Sn (s)} + \text{CuCl}_2 \text{ (aq)}$

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